

Name: \_\_\_\_\_

**IONIC COMPOUNDS HOMEWORK II:** The metals in the following compounds form ions with different charges, so the charge cannot be predicted from the periodic table. Determine the charge on the metal ion and write the name for each compound.

1. SnO
  2. SnBr<sub>4</sub>
  3. FeS
  4. CuSO<sub>4</sub>
  5. ZnCl<sub>2</sub>
  6. CrBr<sub>3</sub>
  7. CoCO<sub>3</sub>
  8. Mn(NO<sub>3</sub>)<sub>2</sub>
  9. FeO
  10. CuCl<sub>2</sub>
  11. Fe<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>
  12. Hg(IO<sub>3</sub>)<sub>2</sub>
  13. Ni<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>
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Write the formula for each of the following compounds

1. tin(IV)chloride
2. iron(III)sulfide
3. mercury(II)oxide
4. cobalt(III)oxide
5. copper(II)sulfate
6. chromium(III)chloride
7. copper(I)phosphate
8. lead(IV)oxide
9. cobalt(III)oxide
10. lead(II)chromate
11. bismuth(III)iodate
12. cobalt(III)hydroxide