

Name: _____

IONIC COMPOUNDS HOMEWORK II: The metals in the following compounds form ions with different charges, so the charge cannot be predicted from the periodic table. Determine the charge on the metal ion and write the name for each compound.

1. SnO tin (II) oxide
 2. SnBr₄ tin (IV) bromide
 3. FeS iron (II) sulfide
 4. CuSO₄ copper (II) sulfate
 5. ZnCl₂ zinc chloride
 6. CrBr₃ chromium (III) bromide
 7. CoCO₃ cobalt (II) carbonate
 8. Mn(NO₃)₂ manganese (II) nitrate
 9. FeO iron (II) oxide
 10. CuCl₂ copper (II) chloride
 11. Fe₃(PO₄)₂ iron (II) phosphate
 12. Hg(IO₃)₂ mercury (II) iodate
 13. Ni₂(SO₄)₃ nickel (III) sulfate
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Write the formula for each of the following compounds

1. tin(IV)chloride SnCl₄
2. iron(III)sulfide Fe₂S₃
3. mercury(II)oxide HgO
4. cobalt(III)oxide Co₂O₃
5. copper(II)sulfate CuSO₄
6. chromium(III)chloride CrCl₃
7. copper(II)phosphate Cu₃(PO₄)₂
8. lead(IV)oxide PbO₂
9. cobalt(III)oxide Co₂O₃
10. lead(II)chromate PbCrO₄
11. bismuth(III)iodate Bi(IO₃)₃
12. cobalt(III)hydroxide Co(OH)₃