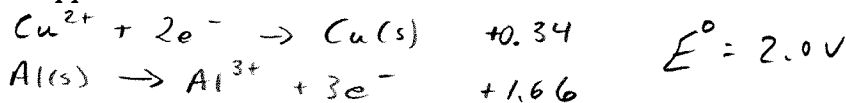


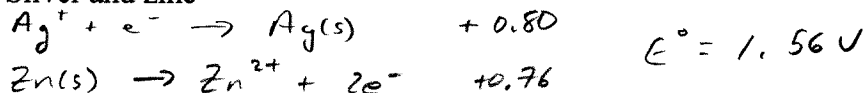
Electrochemistry Worksheet 1 (Voltaic Cells)

1. Write the appropriate half reactions and calculate the cell potential for each of the following electrochemical cells:

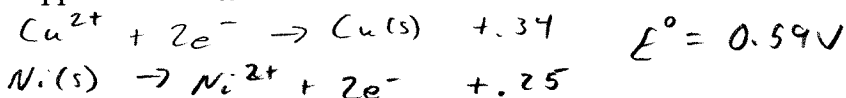
(a) Copper and aluminum



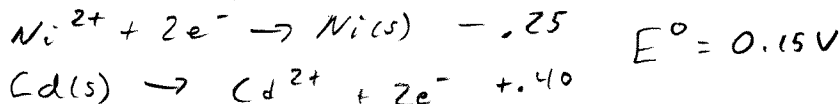
(b) Silver and zinc



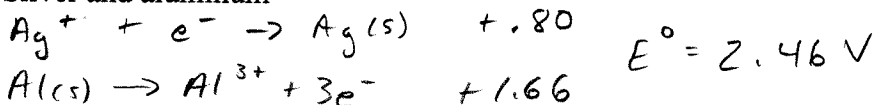
(c) Copper and nickel



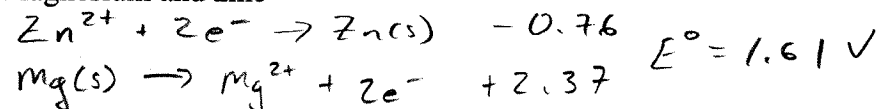
(d) Nickel and cadmium



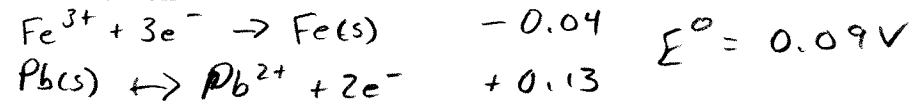
(e) Silver and aluminum



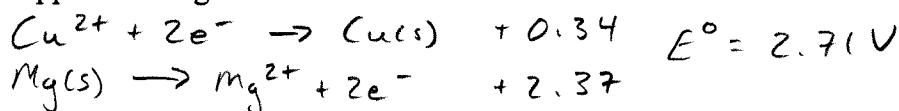
(f) Magnesium and zinc



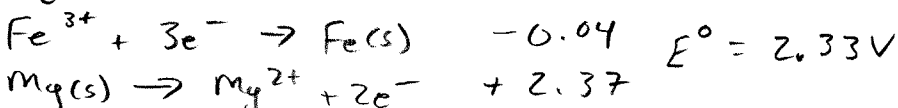
(g) Lead and iron



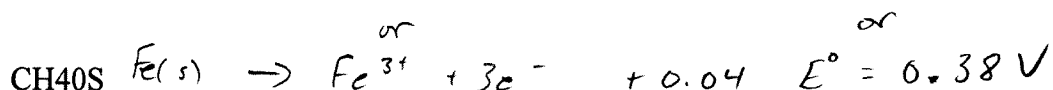
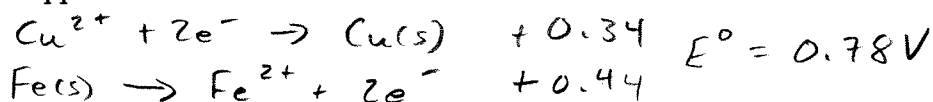
(h) Copper and magnesium



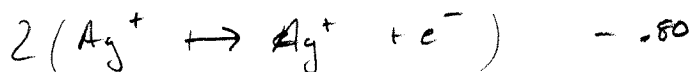
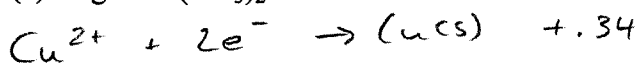
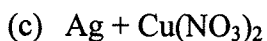
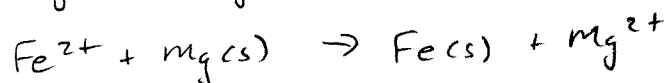
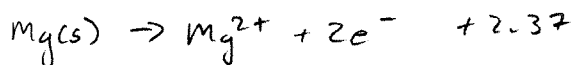
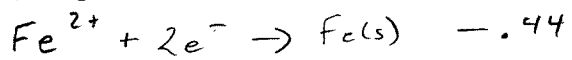
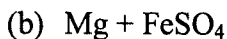
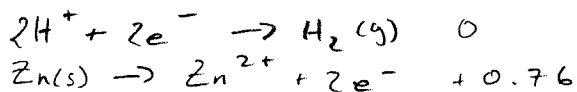
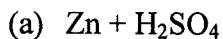
(i) Magnesium and iron



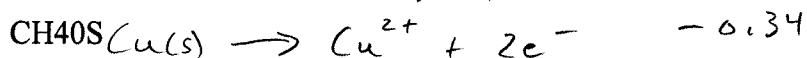
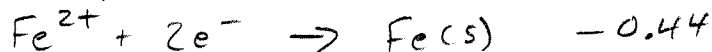
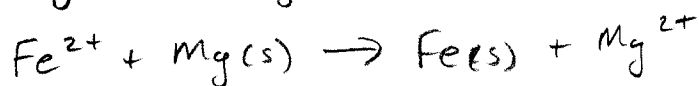
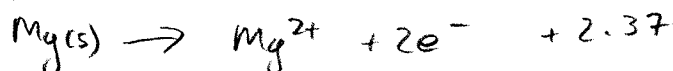
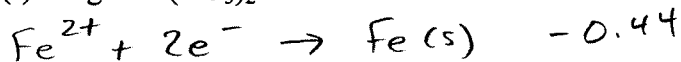
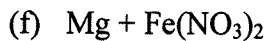
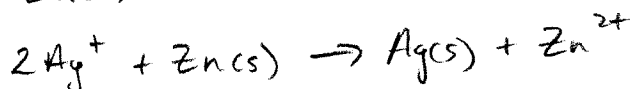
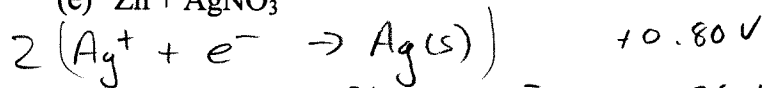
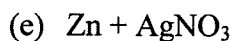
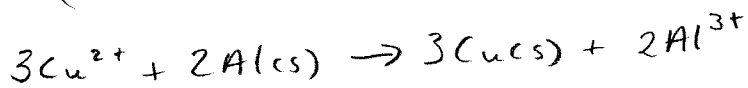
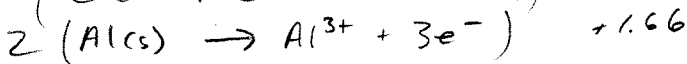
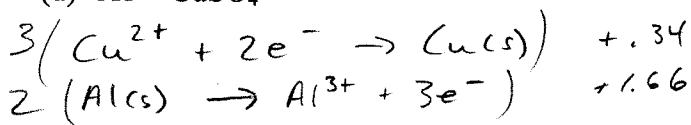
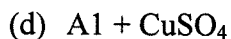
(j) Copper and iron



2. Use standard reduction potentials to determine which of the following reactions can occur. For those that occur write an appropriate balanced chemical reaction.



No Reaction



No Reaction