Solubility Product Worksheet #1

1. Write the expression for the solubility product constant for $SrSO_4$.

2. Write the expression for the solubility product constant for $Al_2(SO_4)_3$.

3. A sample of $Ba(OH)_2(s)$ is added to pure water and allowed to come to equilibrium at 25^oC. The concentration of Ba^{2+} is found to be 0.108 mol/L and that of OH^- 0.216 mol/L. What is the value of K_{sp} for $Ba(OH)_2(s)$?

4. What is the molar solubility of a saturated solution of AgCl? $K_{sp} = 1.6 \times 10^{-10}$

5. What will be the equilibrium concentrations of Ca^{2+} and OH^- in a saturated solution of $Ca(OH)_2$, if its K_{sp} value is 1.3×10^{-6}?

6. Calculate the molar solubility of $Ca(IO_3)_2$ in 0.060 mol/L $NaIO_3$. The K_{sp} of $Ca(IO_3)_2$ is 7.1×10^{-7}.

CH40S